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• **Velten, Ulf, Dr.**

**8046 Zürich (CH)**

• **Schober, Irene, Dr.**

**8008 Zürich (CH)**

• **Bürge, Theodor A.**

**8954 Geroldswil (CH)**

(71) Applicant: **Sika AG, vorm. Kaspar Winkler & Co.  
CH-8048 Zürich (CH)**

(74) Representative: **Blum, Rudolf Emil Ernst et al**

**c/o E. Blum & Co**

**Patentanwälte**

**Vorderberg 11**

**8044 Zürich (CH)**

(72) Inventors:

• **Widmer, Jürg, Dr.**

**8047 Zürich (DE)**

• **Sulser, Ueli**

**8102 Oberengstringen (CH)**

(54) **Cement dispersing polymers for high flow, high strength and selfcompacting concrete**

(57) A multipurpose cement dispersing, shrinkage compensating and anticorrosive polymer and an admixture comprising said polymer, which improve the workability and the early strength development of fresh concrete, are described.

Said polymer is an ester- and amid group-modified acrylic polymer, obtainable by aqueous solution radical copolymerization of (1) acrylic or methacrylic acid with (2) an acrylic or methacrylic methoxypolyalkyleneglycolester monomer having 2 to 300 moles of an oxyalkylene group each having 2 to 3 carbon atoms,

(3) optionally an acrylic or methacrylic methoxypolyalkyleneglycolamide monomer having 2 to 300 moles of an oxyalkylene group each having 2 to 3 carbon atoms,

(4) an acrylic or methacrylic ester of a tertiary aminoalkanol and

(5) optionally an acrylic amide of a primary or secondary aliphatic, cycloaliphatic or aromatic amine.

**Description****Background of the Invention**

- 5   **[0001]** The quality of fresh concrete is determined by the ease and homogeneity with which it can be mixed, transported, compacted and finished. It can also be defined as the amount of internal work, necessary to achieve full compaction.
- [0002]** The rheological behavior of concrete is related to rheological terms such as plasticity and viscoelasticity of its cement paste.
- 10   **[0003]** Cement compositions, such as mortar or concrete, lose their flowability with the progression of time due to a fast stiffening effect caused by an early hydration reaction of cement and water. This loss of flow reduces its workability.
- [0004]** As the workability depends on the conditions of placement, the intended use will determine whether the concrete has the required workability.
- [0005]** Concrete from ready mix plants or mixed on job sites, used in civil engineering constructions, e.g. anchorage of big bridges, base plates or side walls and box culverts, in building structures such as heavy reinforced structures, concrete filling pipe structures or other complicated structures, demands to be fully compacted to achieve its required strength and durability. The existing and conventional method for compaction is by vibration of the freshly placed concrete.
- 15   **[0006]** A new production system for in situ-casted concrete is needed to improve significantly the cost situation as well as the health and safety aspects on the construction site.
- [0007]** Additionally, self compacting concrete leads to a higher productivity, shorter building times and to an improved labor environment.
- [0008]** Increased fluidity, defined e.g. as "flow-table spread", "slump" and "slump-flow", can be effected by using large dosages of water in the concrete, but it is well known that the resulting cement-based structure exhibits insufficient compactness due to excessive bleeding and segregation and consequently will have poor final compressive strength.
- 25   **[0009]** In order to avoid excess amount of water, flowing concrete can be produced by the addition of so called superplasticizers or high range water-reducing admixtures (HRWRs) like sulfonated melamine-formaldehyde polycondensates or naphthalene-formaldehyde polycondensates or ligninsulfonate based admixtures. All of these well known polymers are not capable of causing the treated cement compositions to retain high flowability over a sustained period of time (known as "slump life") without imparting a significant delay in the initial set time and considerable retardation of early strengths. An additional disadvantage is the inconstant and very low flow rate of freshly prepared concrete containing high quantities (e.g. 500 to 700 kg/m<sup>3</sup>) of cement and up to 20% of silica fume and fly ash, which flow rate cannot be improved by the use of conventional HRWRs.
- 30   **[0010]** For heavy reinforced fresh precast concrete it is desired that the cementitious mixture has sufficient fluidity to flow through and around the reinforcement steel-structure to fill out the mold and level off at the top of the mold.
- [0011]** In the past ten years various polymer-admixtures based on so called polycarboxylic acid salts, e.g. copolymers of acrylic acid with acrylic esters of polyalkylene glycols have been proposed for imparting high water reduction, high flow and prolonged slump life to concrete, but most of them do not lead to self compacting concrete without causing a too long retardation of the setting time and the early-strength development.
- 35   **[0012]** Enhancing early, i.e. 1 day, compressive strengths, is of high importance in the precast and prestressed concrete industry, and chemical accelerators such as tertiary alkanolamines, alkali metal and alkaline earth metal thiocyanates, nitrites and halides are known in the art for this purpose and can be added separately to the fresh concrete.
- [0013]** Since calcium chloride and other inorganic salts can initiate corrosion of reinforcing steel embedded in hardened concrete, tertiary alkanolamines which show both, accelerating and corrosion inhibiting properties are the most recommendable chemical accelerators.
- 45   **[0014]** Unfortunately, tertiary amines are well known to act as powerful catalysts for the cleavage of esters and from this reason cannot be premixed with the acrylic ester polymers of the art and stored over longer periods thus causing strongly reduced shelf life of the blend.
- [0015]** In order to improve the stability of such admixtures, a deactivated form of alkanolamines was demanded which is reactivated in an alkaline environment.
- 50   **[0016]** With the introduction of a super high flow- or self compacting concrete, which is or contains a multipurpose polymer simultaneously acting 1) as a HRWR, 2) as set- and strength accelerator and 3) as corrosion inhibitor, these problems could be solved, particularly the need for vibration can be significantly reduced.

55   **Summary of the Invention**

**[0017]** The problem to be solved by the present invention is, that conventional cement dispersing polymers cannot be used as admixtures to produce simultaneously high flow- and high strength selfcompacting concrete.

[0018] The present invention thus provides - as a result of extensive studies - acrylic copolymers, which can solve this problem.

[0019] The invention is based on modified acrylic polymers having a) a poly(oxyalkylene)chain and b) a tertiary alkanolamine group which both are connected to the backbone of the polymer by ester bonds. Said polymers are obtainable by aqueous radicalic copolymerization of acrylic polyalkyleneglycol esters- and -amides with an acrylic acid and an acrylic ester of a tertiary alkanolamine.

[0020] In particular it was found that the relationship between the molar ratio of the free carboxylate groups to the polyoxyethylene ester groups and the polyoxyethylene amide groups in the polymer as well as the length of the polyoxyethylene chain strictly determines the performance of the copolymer as a dispersing, water reducing and slump-keeping agent.

[0021] The accelerating and anticorrosive effect of tertiary alkanolamines in aqueous solutions is known for long time. The inventive polymers react in alkaline cementitious compositions and the alkanolamine group is released from the polymer which is partly adsorbed to the cement surface. Due in particular to their close proximity to the cement surface, the released alkanolamines prove a high set accelerating and shrinkage reducing effect.

[0022] Moreover, results suppose, that the polymers also are adsorbed on the reinforcing steel bars in the fresh concrete where they can act as corrosion inhibitors directly on the surface of the steel.

#### Detailed Description of the Invention

[0023] A water-soluble acrylic copolymer with cement dispersing, set-accelerating and anticorrosive properties to steel, as well as an admixture comprising an amount of from 5 to 95 % of said water soluble acrylic copolymer, are described.

[0024] Said polymer is obtainable by copolymerization of an  $\alpha,\beta$ -olefinic monocarboxylic acid or its salt shown by formula 1 given below, a second acrylic monomer shown by formula 2 given below, optionally a third acrylic monomer shown by formula 3 given below, a fourth acrylic monomer shown by formula 4 given below and optionally a fifth acrylic monomer given by formula 5,

such that the molar ratio of constituent monomer units 1, 2, 3, 4, 5 is

$$a : b : c : d : e = (0.1-0.9) : (0.01-0.80) : (0-0.80) : (0.001-0.3) : (0-0.5),$$

and

$$a : (b+c+d+e) = 0.1 : 0.9 \text{ to } 0.9 : 0.1, \text{ whereby}$$

a designates the molar concentration range of constituent monomers 1,

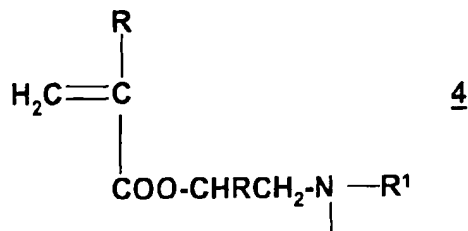
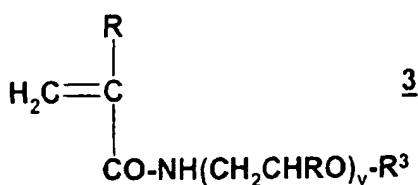
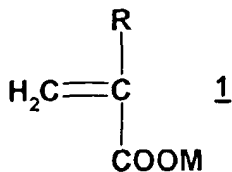
b designates the molar concentration range of constituent monomers 2,

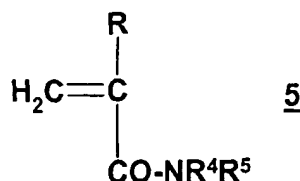
c designates the molar concentration range of constituent monomers 3,

d designates the molar concentration range of constituent monomers 4,

e designates the molar concentration range of constituent monomers 5

said monomers having the structural formulas shown below:





wherein each R independently from each other represents a hydrogen atom or a methyl group, R<sup>1</sup> and R<sup>2</sup> independently from each other represent a C<sub>1</sub> - C<sub>4</sub>- alkyl rest, a hydroxyethyl(CH<sub>2</sub>CH<sub>2</sub>OH) group, an acetoxyethyl group(CH<sub>2</sub>CH<sub>2</sub>OCOCH<sub>3</sub>) group, hydroxy-isopropyl(CH<sub>2</sub>-CHOH-CH<sub>3</sub>) group, an acetoxy-isopropyl (CH<sub>2</sub>-CHOCOCH<sub>3</sub>) group, or R<sup>1</sup> and R<sup>2</sup> may together with the nitrogen atom to which they are bound, form a morpholine ring;

R<sup>3</sup> is a C<sub>1</sub>-C<sub>4</sub>-alkyl rest,

R<sup>4</sup> and R<sup>5</sup> independently represent an aliphatic, cycloaliphatic, araliphatic or aromatic rest or R<sup>4</sup> and R<sup>5</sup> may together with the nitrogen atom to which they are bound form a morpholine- or imidazole ring system,

M represents hydrogen, an alkali metal, an alkaline earth metal, an aluminum ion, an ammonium- or organic ammonium group, such as an ammonium group derived from alkylamines or alkanolamines, in particular amines comprising C<sub>1</sub> to C<sub>4</sub> alkyl and/or C<sub>1</sub> to C<sub>4</sub> alkanol groups,

x and y independently represent integers from 2-300.

**[0025]** In the solvent polymerization using an aqueous solvent, a water soluble polymerization initiator such as ammonium salts, alkali metal salts of a peroxy acid or hydrogen peroxide is used as the polymerization initiator.

**[0026]** Further, to control the molecular weight, a chain transfer agent such as alkali sulfite, hydrogensulfite, mercaptoethanol, thioglycolic acid or thioglycerol(3-mercapto-1,2-propanediol) may be used.

**[0027]** It is preferable that the inventive polymer has a number- average molecular weight of 1000 to 100000, still more preferably 1000 to 30000.

**[0028]** When the molecular weight is too large, the resulting copolymer will exhibit a poor dispersing effect, while, when it is too small, the resulting copolymer will exhibit not only poor dispersity but also a poor slump retaining effect. Moreover, a defined molar proportion of 1-5 to 1 of carboxylic-groups to methoxypolyalkyleneglycol ester side chains and methoxypolyalkyleneglycol amide side chains at the backbone of the copolymer is preferred.

**[0029]** The average molecular weight is determined by gel permeation chromatography using a set of defined polyalkyleneglycols as standard reference

**[0030]** Following a further aspect of the present invention the inventive copolymers are employed for preparing a cement-dispersing, set accelerating admixture having shrinkage-reducing and anticorrosive properties to steel. Hence, further to the invention, said admixture preferably contains at least one defoaming and/or at least one air-controlling agent.

**[0031]** Still a further aspect of the present invention is a mortar, concrete, cement or cementitious binder containing the modified acrylic polymer according to the present invention in an amount of 0.01 to 10% by weight of the binder, said mortar or concrete having a unit content of binder composed of cement or a mixture of cement and latent hydraulic or inert microscopic powder of 150 to 800 kg/m<sup>3</sup>, preferably of 250 to 650 kg/m<sup>3</sup>.

**[0032]** In a preferred embodiment, the cement is Portland cement, white cement, high alumina cement, and cement blends wherein the latent hydraulic or inert microscopic powder is fly ash, slag, natural pozzolane, silica fume, burnt oil shale, metakaolin or calcium carbonate.

**[0033]** The inventive admixture can be used in liquid form, e.g. by addition of water, in powder form and can be added before, during or after the grinding operation of the cement or the cementitious binder.

## EXAMPLES

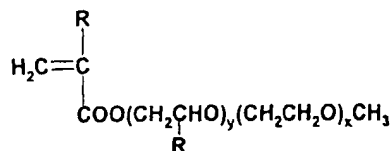
**[0034]** In the following, some copolymers of the present invention and cement dispersing admixtures comprising such polymers, the method for producing the same and the cement containing composition of the present invention are explained in more detail with the following examples E1 - E6 and comparative examples C1, C2, and C3. Moreover, for comparison, commercially available polymers C7, C8 and C9 are also used for the preparation and testing of flowing concrete (test example 1) and high-strength, self compacting concrete of low water-to-cement ratio (test examples 2 and 3).

**[0035]** In test examples 4-6 the corrosion inhibiting effect of the inventive polymers on mild steel and on reinforcing steel is demonstrated and compared to comparative examples.

[0036] However, it must be noted that all these examples are given for illustrative purposes only and are not supposed to limit the invention, as defined by the claims, appended hereto.

[0037] The following polyalkyleneglycol-based acrylic esters and -amides have been used in the examples described below (table 1):

Table 1

**Acrylic ester monomers:****Monomer 2-1:**

Methoxypolyoxyalkylene acrylate

x=22, y=0

**Monomer 2-2:**

Methoxypolyoxyalkylene (meth)acrylate

x=22, y=0

**Monomer 2-3:**

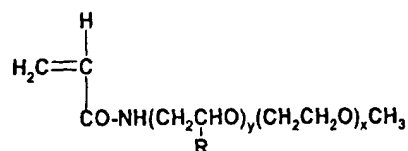
Methoxypolyoxyalkylene (meth)acrylate

x=44, y=0

**Monomer 2-4:**

Methoxypolyoxyalkylene(methacrylate

x=16, y=5

**Acrylamide-Monomers:****Monomer 3-1:**

Methoxypolyoxyalkylene acrylamide

x=11, y=0

**Monomer 3-2:**

Methoxypolyoxyalkylene acrylamide

x=19, y=3

**Monomer 3-3:**

Methoxypolyoxyalkylene acrylamide

x=32, y=10

**Monomer 3-4:**

Methoxypolyoxyalkylene acrylamide

x=6, y=29

**Monomer 3-5:**

Methoxypolyoxyalkylene acrylamide

x=16, y=5

Comparative Example C1 (Polymer C1)

[0038] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser, N<sub>2</sub> -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1) 493 g (0.24 mol) of monomer 3-3, 82 g (0.04 mol) of monomer 3-4 and 6 g of 50%-sodium hydroxide in 180g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of M<sub>n</sub> = 9300 g/mol was obtained.

Example E 1 (Polymer E1)

[0039] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser, N<sub>2</sub> -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 493 g (0.24 mol) of monomer 3-3, 82 g (0.04 mol) of monomer 3-4, 15 g (0.05 mol) dimethylaminoethyl acrylate and 6 g of 50%-

sodium hydroxide in 180 g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n$  = 9900 g/mol was obtained.

#### Comparative Example C2 (Polymer C2)

**[0040]** Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72g (1mol) of freshly distilled acrylic acid (monomer 1), 513.5 g (0.25 mol) of monomer 2-2, 50 g of polypropyleneglycol of MW 4000 and 6 g of 50%-sodium hydroxide in 180 g Water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water.

**[0041]** After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n$  = 10100 g/mol was obtained.

#### Example E2 (Polymer E2)

**[0042]** Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 513.5 g of monomer 2-2, 59 g of polypropyleneglycol of MW 4000, 2.15 g (0.015 mol) of dimethylaminoethyl acrylate and 6 g of 50%-sodium hydroxide in 200 g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight  $M_n$  = 9900 g/mol was obtained.

#### Comparative Example C3 (Polymer C3)

**[0043]** Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 685 g (0.65 mol) of monomer 2-1, 47 g (0.085 mol) of monomer 3-1, and 6g of 50%-sodium hydroxide in 250 g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n$  = 11100 g/mol was obtained.

#### Example E3 (Polymer E3)

**[0044]** Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 685 g (0.65 mol) of monomer 2-1, 47 g (0.085 mol) of dimethylaminoethyl acrylate and 6 g of 50%-sodium hydroxide in 250 g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with

a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n = 10700$  g/mol was obtained.

#### 5 Comparative Example C4 (Polymer C4)

[0045] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1) 516 g (0.49 mol) of monomer 2-1, 95 g (0.09 mol) of monomer 3-2 and 6g of 50%-sodium hydroxide in 200g Water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n = 10200$  g/mol was obtained.

#### Example E4 (Polymer E4)

[0046] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 516 g (0.49 mol) of monomer 2-1, 95 g (0.09 mol) of monomer 3-2 and 2.86 g (0.02mol) of dimethylaminoethyl acrylate in 200g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight  $M_n = 9700$  g/mol was obtained.

#### Comparative Example C5 (Polymer C5)

[0047] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1) 580 g (0.55 mol) of monomer 2-1, 28 g (0.12 mol) of N,N-dicyclohexyl acrylamide and 6 g of 50%-sodium hydroxide in 220 g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n = 10300$  g/mol was obtained.

#### Example E5 (Polymer E5)

[0048] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser,  $N_2$ -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C. Then simultaneously were added in 60 minutes, a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 580 g(0.55 mol) of monomer 2-1, 2.86 g (0.02 mol) of dimethylaminoethyl acrylate and 6g of 50%-sodium hydroxide in 220 g water, b) a solution of 15 g of 35%-hydrogen peroxide and c) 4g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight of  $M_n = 9500$  g/mol was obtained.

Example E6(Polymer E6)

[0049] Into a 2 liter-glass reactor, equipped with a stirrer, reflux condenser, N<sub>2</sub> -inlet tube and automatic dosing mechanism, were placed 0.30 g of ferrous sulfate and 5 g of thioglycerol dissolved in 380 g water. The inside atmosphere of the flask was replaced with stirring by nitrogen gas and the solution heated to 60°C.

[0050] Then simultaneously were added in 60 minutes a) an aqueous solution of 72 g (1 mol) of freshly distilled acrylic acid (monomer 1), 317 g(0.19 mol) of monomer 2-4,

317 g (0.19 mol) of monomer 3-5 and 2.15 g of of dimethylaminoethyl acrylate in 200g water,

b) a solution of 15 g of 35%-hydrogen peroxide and c) 4 g ascorbic acid in 25 g water. After the completion of the addition, in course of which the temperature raised to 75°C, the reaction system was maintained at that temperature for 15 minutes. Finally, the mixture was cooled to 25°C and 2 g of methoxy-polyoxyalkyleneamine of MW=2000 with a PO/EO mol. ratio of 29:6 were added. 35 g of 50%-sodium hydroxide were then added dropwise to adjust the pH-value to 5.0. Water was added to give a 40%-solution. A polymer of a number average molecular weight  $M_n=9900g/mol$  was obtained.

Comparative Example C6 (Polyacrylic acid)

[0051] SOKALAN PA 25 CL PN (BASF, Badische Anilin& Sodafabrik) a 50% aq. solution of a partially neutralized polyacrylic acid of weight-average molecular weight 4000 was used as a comparative dispersant for concrete.

Comparative Example C7 (Melamine polymer)

[0052] Melment F-10 (Süddeutsche Kalkstick-stoffwerke,Trostberg, BRD), a commercial dispersing agent for hydraulic cement compositions, is the sodium salt of a sulfonated melamine formaldehyde polycondensate of molecular weight of about 15'000.

Comparative Example C8

[0053] MIGHTY-150 (KAO Corp.,Tokyo), a commercial dispersing agent for hydraulic cement masses, is the sodium salt of a sulfonated naphthalene-formaldehyde polycondensate of average molecular weight of about 5'000-6'000.

TEST-EXAMPLES

[0054] These examples were performed to demonstrate the improved fluidizing effects of the inventive polymers on fresh concrete and their accelerating action on the development of the early(1d) compressive strengths. The inventive polymers E1 to E5, were tested as fluidizers in flowing concrete (test-example 1) and as admixtures to improve the flowability and slump life of high -strength concrete of low water-to-cement ratio and high binder(cement + silica-fume) content(test example 2).

[0055] Those comparative polymers(C1 to C8) described above were also tested and compared in this context.

Test-Example 1Flowing concrete

[0056] Use of inventive polymers and comparative polymers for flowing concrete.

[0057] The consistency of freshly prepared concrete i.e. the mobility or viscosity , is the most important characteristic of workability. For measuring the consistency of concrete a "flow table spread" according to DIN 1048, part I is used in industry.

[0058] Sometimes the "slump test" according to ASTM C143 is additionally used.

[0059] For purposes of this experiment the flow table spread was determined by placing concrete in an iron form on a two-part table (70 x 70 cm).By removing the form, a concrete body having a truncated cone shape is prepared. Then, the are of the table is lifted on one side for 4cm, and allowed to fall. This procedure is carried out 15 times , and the concrete spreads. The average diameter of the formed cake corresponds to the flow table spread.

[0060] For the slump test, three layers of concrete are put into a mold having a shape of a truncated cone and having certain dimensions, and compressed with 25 pushes of an iron bar. At the top, the concrete is stripped off evenly, then, the form is vertically removed. The concrete body will sink in by itself. The slump is measured by determining the vertical difference between the top of the mold and the displaced original center of the top surface of the test sample.

[0061] In order to compare the obtained test results and to bring them into a relation with the consistency, the freshly



prepared concrete (see DIN 18555, Part2) may be divided into consistency ranges:

Table 2:

Consistency Ranges of Freshly Prepared Concrete		
Denotation	Flow Table Spread (cm)	Slump (cm)
K1= rigid	< 30	< 1
K2= semi-plastic	30 to 40	1 to 9
K3= plastic	41 to 50	10 to 15
K4= flowing concr.	> 50	> 15

**[0062]** Fluidizers are used when specific construction applications are necessary. Flowing concrete is used when high inserting rates (e.g., from 50 to 150 m<sup>3</sup>/hour) are required, or when the form and reinforcement of a construction part do not allow a compacting of the concrete by vibration.

**[0063]** Concretes having K2 or K3 consistencies may be prepared from a concrete of K1 consistency by adding fluidizers(also designated as superplasticizers) when increased mechanical strength at an equal remaining workability shall be obtained.

**[0064]** For a freshly prepared concrete, the fluidizing effect is dependent on the dosage of the superplasticizer. Usually, from 0.2 to 1.5% solid matter quantities(in dissolved form),referred to the weight of cement, are added.

**[0065]** To a high degree, the effect is also dependent on the chemical structure and the molecular weight of the polymer, which forms the basis of the fluidizer.

**[0066]** In order to demonstrate the increased effectiveness of the inventive copolymers, the flow behavior of concrete mixtures containing the copolymers E1 to E5 was measured according to DIN 1048, part1, and ASTM C143. As a comparison, the polymers C-1 to C-8 were also tested with the same methods.

Table 3:

Composition of the fresh concrete mixtures	
Components:	Quantity in kg
Normal Portland Cement.TypeI	7.5
Netstal filler (chalk filler)	1.5
Rhine sand "Epple" up to 1.2 mm.*	9.5
Rhine sand "Epple" 1.2 to 4.0 mm.*	8.0
Rhine sand "Epple" 4.0 to 8.0 mm.*	4.5
Mine gravel 8 to 16 mm.*	9.5
Mine gravel 16 to 32 mm.*	17.0
Total Water, including mixing water and water of the copolymer solution	3.45,
Copolymer(solid) or comparative polymer, used as fluidizer	0.04, dissolved in the mixing water

\* washed and dried

#### Preparation and handling of the concrete specimen

**[0067]** The cement and the aggregates were premixed for 15 seconds in a 50 liter forced circulation mixer for concrete. The mixing water, containing the fluidizer, was added slowly under stirring over 20 seconds. The batch was then mixed in a wet state for an additional 60 seconds. A part of the fresh concrete was then immediately filled into the mold for the determination of the flow table spread and the slump.

**[0068]** Immediately after measuring the flow table spread, test bodies having edges of 12 x 12 cm were prepared, and the compressive strength was measured after 1, 7, and 28 days in accordance with DIN 1048, Part 1. The determination of initial setting was carried out according to ASTM-C 403.

**[0069]** Additionally, the copolymers of the present invention were compared to the comparative polymers C-1 to C-3.

**[0070]** As mentioned above, the flow table spread and slump was measured immediately after mixing, and remeasured at 60 and 120 minutes after mixing. A mixing up of the concrete for five seconds was carried out before each new measurement.

[0071] Concrete test mixtures No 1 to 11 which were prepared under the same conditions, were then subjected to the above-described examination of flow table spread and slump depending on the time.

[0072] The results are summarized in Table 4. They show a high water reduction and a surprisingly long lasting constancy of the flow table spread and the slump of up to 120 minutes in test mixtures nos. 2, 4, 6, 8, 10, 11 containing the copolymers according to the invention. From a comparison of these mixtures with comparative mixtures nos. 1, 3, 5, 7 and 9 containing the alkanolaminefree polymers, it can be seen, that the comparative test mixtures show a strong retardation of the early strength development. Also, comparative mixtures no 12, 13 and 14 containing polyacrylic acid and melamine- and naphthalene polycondensate show a considerable stiffening tendency already 60 minutes after mixing.

**Table 4** Time-dependent flow behavior and compressive strengths of flowing concrete of water-to-cement ratio  $W/C = 0.46$  using inventive and comparative polymers as fluidizers.

Test-Mix- ture No	Poly- mer exam- ple	Dosage in % of cement weight	Flow table spread /slump in cm, x minutes after mixing			Compressive strength in New- ton/mm <sup>2</sup> y days after mixing			% -Air con- tent after mix
			x=0	60	120	y=1	7	28	
1	C1	0.3	61/24	60/24	50/19	14.0	42.8	47.9	1.9
2	E1	0.3	59/22	57/21	51/17	20.4	43.3	48.5	2.0
3	C2	0.3	52/19	53/20	50/18	17.5	39.7	48.0	2.0
4	E2	0.3	50/18	52/19	49/19	23.2	44.3	46.0	1.8
5	C3	0.3	56/21	54/22	50/20	14.9	39.0	47.1	1.9
6	E3	0.3	55/20	54/21	52/19	20.4	42.3	50.2	1.6
7	C4	0.3	54/20	52/20	50/20	15.6	41.2	50.0	2.7
8	E4	0.3	60/22	60/24	53/20	22.3	44.5	50.2	1.8
9	C5	0.3	56/22	55/19	48/18	17.0	45.0	50.3	2.5
10	E5	0.3	61/20	60/25	52/20	20.5	42.0	49.8	1.8
11	E6	0.3	60/21	61/23	61/22	21.9	43.0	51.2	2.0
12	C6	0.3	30/0	-	-	-	22.4	45	2.8
13	C7	0.3	42/12	-	-	26.9	48.1	60.9	1.5
14	C8	0.3	45/14	33/1	-	21.0	43.5	50.0	1.9

[0073] The measurement of flowing properties of fresh mixtures of high flowing -high strength concrete of very low water-to-cement ratio( $W/C$ ) is described in the next test-example.

#### Test-Example 2

##### High flowing-high strength concrete.

[0074] High flowing-high strength concrete of very low water-to-cement ratio and very high content of binder(cement + silicafume) is increasingly demanded by the building- and construction industry. Preparation and testing are prescribed in Japanese Industrial Standards(JIS-A).

##### Preparation of the concrete mixtures

[0075] In a mixing ratio as shown in table 5, ordinary Portland cement, silica fume, fine aggregates and coarse aggregates (gravel) were sequentially placed inside a forced mixing-type mixer of 50 liters volume. The cement and the aggregates were premixed for 15 seconds, and then the mixing water, containing the fluidizer and 0.02 % (related to the weight of fluidizer) of a synthetic air detrainer, was added slowly under stirring over 20 seconds. The batch was then mixed in a wet state for 3 minutes. After mixing, the mixture was transferred to a mixing boat and retempering

was conducted at a predetermined number of times every 60 minutes and the slump flow and slump with the progression of time was measured for up to 120 minutes according to JIS-A 1101. The procedures specified in JIS-A 1123 and JIS-A 6204 were employed to measure air content and time-dependent compressive strengths.

**Table 5: Concrete Mix Proportion for Self Compacting Concrete**

			UNIT CONTENT (kg/m <sup>3</sup> )				
W/B	S/A	Polymer-Dosage	W	C	SF	S	G
22%	39%	1.6%	165	675	75	600	950

**Raw materials:**

**W** = Mixing water, including water of the added fluidizer.

**C** = Cement: Normal Portland Type 1

**SF** = Undensified Silicafume: density: 2.2-2.7, surface: 100.000-250.000 cm<sup>2</sup>/g

**S** = Sand Eppl: 0-1.2mm, 300kg  
1.2-4mm, 300kg

**G** = Gravel Eppl: 4-8mm, 280kg  
8-16mm 670kg

**A** = Aggregate: Sand + Gravel

**B** = Cement + Silicafume

**Polymer-dosage:** dosage (weight%) of solid polymer, contained in the fluidizer, in relation to the cement weight.

Results of the evaluation of the mixtures, containing inventive and comparative polymers, are shown in table 6.

**Table 6:** Time-dependent flow behavior and compressive strengths of Flowing concrete of water-to-cement ratio  $W/C=0.22$ , using inventive and comparative polymers as fluidizers.

Test-Mixture No	Polymer designation	Polymer dosage in %	Slump-flow in cm x minutes after mixing			Compressive strength in Newton/mm <sup>2</sup> y days after mixing		% Air content after mixing
			x=0	60	120	y=1	28	
1	C1	1.6	51	58	60	11.7	97.9	2.1
2	E1	1.6	50	52	60	22.0	101.0	2.0
3	C2	1.6	51	55	59	13.7	102.0	2.4
4	E2	1.6	56	61	59	29.1	100.0	2.0
5	C3	1.6	54	53	51	8.8	103.4	1.9
6	E3	1.6	61	60	59	22.6	105.0	1.8
7	C4	1.6	60	61	60	14.6	100.0	2.1
8	E4	1.6	60	64	62	24.7	102.5	2.2
9	C5	1.6	54	58	60	16.0	101.1	2.4
10	E5	1.6	57	61	59	23.9	104.3	2.3
11	E6	1.6	60	62	62	24.0	105.0	2.4
12	C6	-	<35	-	-	-	63.9	8.4
13	C7	1.6	43	-	-	31.1	102.4	1.8
14	C8	1.6	48			19.3	93.2	3.8

[0076] It is evident from table 6, that the test mixtures 2,4,6,8 10 and 11, containing the inventive, dimethylaminoethyl acrylate-modified polymers, show an improved setting behavior and early-strength development as compared to comparative mixtures 1,3,5,7 and 9 which exhibit strong retardation of initial setting and early strengths whereas test mixtures 12, 13 and 14, which contain conventional polymers, show only poor flowability.

### Test-Example 3

#### Corrosion inhibition of mild steel

[0077] The anticorrosive activity of the polymers was tested by the stacked steel test, as is known.

[0078] An aqueous solution containing 1.6% by weight of the polymers and having a water hardness of 177 ppm of calcium was prepared for use in the test.

[0079] Mild steel cylinders of 3.3 - 4.5 cm length and 2.5 cm in diameter were used as the test material. All surfaces except one end were coated with an epoxy coating.

[0080] The exposed test end of each slug was polished with a power grinding wheel, then was covered with about

lg of the 1.6%-polymer solution. The exposed end of the second steel slug was placed over the test end to spread the film and prevent evaporation. The stacked steel slugs were then placed in an equilibrated desiccator over water. Tests were carried out at 20°C.

[0081] Tap water of 177 ppm hardness was used as a control in each test.

[0082] The test results (table 7) showed, that solutions containing the inventive polymers even at low concentrations were non corrosive whereas none of the comparative polymers did develop an anticorrosive activity.

Table 7

Time-dependent corrosion of mild steel in water: Inhibition by inventive polymers				
Test No.	Polymer designation	Concentration of the aq. polymer solution	Period of examination [days]	Condition of steel after test
1	Control	-	1d	K
2	Control	-	7d	K-st
3	C1	1.6%	1d	Inh
4	C1	1.6%	7d	K
5	E1	1.6%	1d	Inh
6	E1	1.6%	7d	Inh
7	C2	1.6%	7d	K
8	E2	1.6%	7d	Inh
9	E3	1.6%	7d	Inh
10	C3	1.6%	1d	K
11	C4	1.6%	1d	K
12	E4	1.6%	7d	Inh
13	E4	1.6%	28d	Inh
14	E5	1.6%	7d	Inh
15	E6	1.6	7d	Inh
16	C5	1.6%	1d	K

K: Distinct traces of corrosion visible on the steel

K-st: Strong corrosion visible

Inh: No corrosion visible

#### Test-Example 4

**Measurement of the anticorrosive effect of inventive polymers on mild steel in CaCl-solutions by cyclic voltammetry(CV)**

[0083] In this example, the influence of the inventive, alkanolamine-modified corrosion inhibiting polymers in comparison to non modified polymers C1,C2 and C3 on mild steel specimens was investigated.

[0084] The examination was carried out by means of potentiodynamic polarization measurements, performed in aqueous, chloride containing solutions using a potentiostat/galvanostat with scan option and a Camec II station.

#### Test conditions:

[0085]

$$U_{\max} \pm 5V, v_u = 1-8350\text{mV/min}$$

- working electrodes: polished steel plates (mild steel ST37)
- $v = 10\text{mV/s}$

- concentration of inventive and comparative polymers: 1.60% based on solids.

[0086] The test solution was neither stirred, nor degassed. The electrode was conditioned for one hour at  $i = 0.0 \text{ mA/cm}^2$ , then polarization was started.

5 [0087] Test period: 1 hour;  $i_{\text{max}} = 0.5 \text{ mA/m}^2$   $j_0$  and  $j_u$  depending on

currents ( $i_{\text{max}} = 30\text{-}50 \text{ mA}$ )

10 [0088] After a further conditioning at  $i = 0.50 \text{ mA/cm}^2$  for 30 minutes, the CV was performed in cathodic direction during 30 minutes.

[0089] As can be seen from the data in table 8, all inventive polymers perform substantially better than the comparative polymers as well as the control, showing higher positive values of both, the pitting- and the repassivation potentials.

Table 8:

15

Cyclic voltammetry at mild steel electrode in saturated  $\text{Ca(OH)}_2 + 0.09\text{M-Na(OH)}$

Test No.	Polymer designation	Concentration of polymer in test solution [%]	Pitting potential vs SCE Ø 3 scans [mV]	Repassivation potential vs SCE Ø 3 scans [mV]
20	1	Control	---	-520
	2	C1	+235	-575
	3	E1	+1200	+990
25	4	C2	+375	-600
	5	E2	+990	+720
	6	C3	+405	-590
30	7	E3	+1180	+980
	8	C4	+630	-410
	9	E4	+1000	+720

### 35 Test-Example 5

[0090] In the following, the influence of inventive tert.-alkanolamines containing polymers on the shrinkage behavior of a mortar of high cement content is demonstrated by the measurement of the strain of  $4 \times 4 \times 16 \text{ cm}^3$  prisms fitted with an inserted neck.

40 Composition of the fresh mortar:

[0091]

45

Portland cement type CEM I 42.5	0.750 kg
Sand 0-8 mm	3.150 kg
Water	0.350 kg
Inventive or comparative polymer	0.012 kg

50 [0092] The test specimen were initially stored for 24 hours at  $20^\circ\text{C}$  and 95% RH, subsequently, at  $23^\circ\text{C}$  and 50% RH.

Table 9:

55

Effect of inventive, shrinkage compensating polymers on the strain of mortar prisms.

Test No.	Polymer Designation	Dosage of solid Polymer, calculated on cement weight(%)	Strain after 28 d
1	Control	1.6	-0.057

Table 9: (continued)

Effect of inventive, shrinkage compensating polymers on the strain of mortar prisms.			
Test No.	Polymer Designation	Dosage of solid Polymer, calculated on cement weight(%)	Strain after 28 d
2	C1	1.6	-0.045
3	E1	1.6	-0.019
4	C2	1.6	-0.045
5	E2	1.6	-0.024
6	C3	1.6	-0.049
7	E3	1.6	-0.036
8	C4	1.6	-0.047
9	E4	1.6	-0.022
10	C5	1.6	-0.044
11	E5	1.6	-0.022

[0093] This example clearly demonstrates the shrinkage compensating effect of up to 50% of the inventive polymers on mortar pieces whereas comparative polymers show only poor reduction as compared to the control specimen.

[0094] While there are shown and described presently preferred embodiments of the invention, it is to be distinctly understood that the invention is not limited thereto but may be otherwise variously embodied and practiced within the scope of the following claims.

#### Claims

1. A water-soluble acrylic copolymer with cement dispersing, set-accelerating and anticorrosive properties to steel,

obtainable by copolymerization of an  $\alpha,\beta$ -olefinic monocarboxylic acid or its salt shown by formula 1 given below,

a second acrylic monomer shown by formula 2 given below, optionally a third-acrylic monomer shown by formula 3 given below, a fourth acrylic monomer shown by formula 4 given below and optionally a fifth acrylic monomer given by formula 5,

such that the molar ratio of constituent monomer units 1, 2, 3, 4, and 5 is

$$a : b : c : d : e = 1 : (0.01-0.80) : (0-0.80) : (0.001-0.3) : (0-0.5),$$

and

$$a : (b+c+d+e) = 0.1 : 0.9 \text{ to } 0.9 : 0.1, \text{ wherein}$$

a designates the molar concentration range of constituent monomers 1,

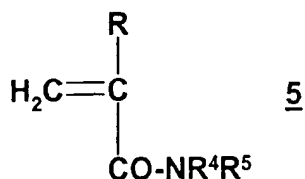
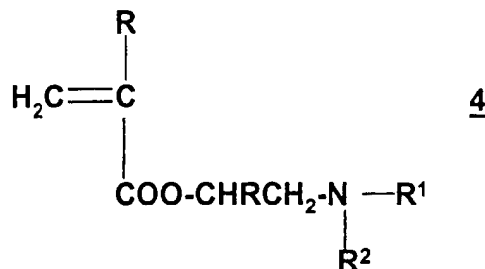
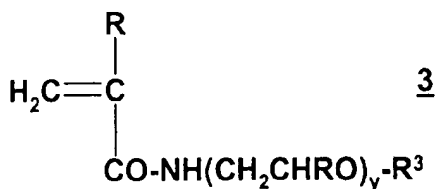
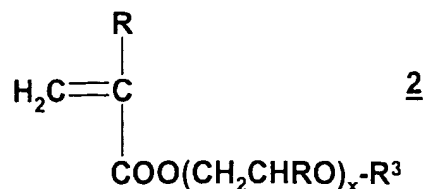
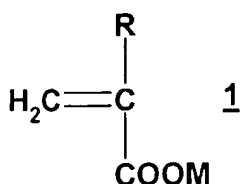
b designates the molar concentration range of constituent monomers 2,

c designates the molar concentration range of constituent monomers 3,

d designates the molar concentration range of constituent monomers 4,

e designates the molar concentration range of constituent monomers 5,

said monomers having the structural formulas shown below:



wherein each R independently from each other represents a hydrogen atom or a methyl group,  
 R<sup>1</sup> and R<sup>2</sup> independently from each other represent from each other a C<sub>1</sub> - C<sub>4</sub>- alkyl rest, a hydroxyethyl  
 (CH<sub>2</sub>CH<sub>2</sub>OH) group, an acetoxyethyl-group(CH<sub>2</sub>CH<sub>2</sub>OCOCH<sub>3</sub>) group, a hydroxy-isopropyl (CH<sub>2</sub>-CHOH-CH<sub>3</sub>)  
 group, an acetoxy-isopropyl(CH<sub>2</sub>-CHOCOCH<sub>3</sub>) group, or R<sup>1</sup> and R<sup>2</sup> may together with the nitrogen atom to  
 which they are bond, form a morpholin ring;

R<sup>3</sup> is a C<sub>1</sub>-C<sub>4</sub>-alkyl rest,

R<sup>4</sup> and R<sup>5</sup> independently represent an aliphatic, cycloaliphatic, araliphatic or aromatic rest or R<sup>4</sup> and R<sup>5</sup> may  
 together with the nitrogen atom to which they are bond form a morpholine- or imidazole ring system,

M represents hydrogen, an alkali metal ion, an alkaline earth metal ion, an aluminum ion, an ammonium- or  
 organic ammonium group;

x and y independently represent integers from 2-300.

2. A modified acrylic polymer according to claim 1, wherein the monomer 1 is acrylic acid or methacrylic acid.

3. A modified acrylic polymer according to claim 1 or 2, wherein the monomer 2 is a methoxy-polyoxyethylene acrylate  
 of a molecular weight of 350 to 5000g/mol.

4. A modified acrylic polymer according to anyone of claims 1 to 3, wherein the monomer 3 is a methoxy-polyoxy-  
 alkylene acrylamide of a molecular weight of 500 to 5000g/mol.

5. A modified acrylic polymer according to anyone of claims 1 to 4, wherein the monomer 4 is the acrylic ester of a  
 tertiary alkanol amine.

6. A modified acrylic polymer according to claim 4, wherein the monomer 4 is dimethylaminoethyl acrylate.

7. A modified acrylic polymer according to anyone of claims 1 to 6, wherein the monomer 5 is the acrylamide of  
 oxazolidine.



8. A modified acrylic polymer according to anyone of claims 1 to 6, wherein the monomer 5 is the acrylamide of dicyclohexylamine.
- 5 9. A modified acrylic polymer according to claims 1 to 8, having a number average molecular weight in the range of 5'000 to 50'000.
- 10 10. A method for preparing a modified acrylic polymer according to claim 1, wherein the monomers 1, 2, 3, 4, and 5 are reacted by aqueous solution radical copolymerization in the presence of aqueous hydrogen peroxide as initiator.
- 10 11. The method of claim 11, wherein an organic thiol compound is used as a chain-transfer agent.
12. The method of claim 1, wherein thioglycerine is used as a chain-transfer agent.
- 15 13. A modified acrylic polymer obtainable by the method of one of claims 10 to 12.
14. A cement dispersing, set accelerating admixture having anticorrosive and preferably also shrinkage reducing properties to steel, that contains a modified acrylic polymer according to anyone of claims 1 to 9 and 13 in amounts of from 5 to 95 %, and that is in liquid or powder form.
- 20 15. The cement admixture of claim 14, containing at least one defoaming or antifoaming agent.
16. The cement admixture of claim 14 or 15, containing at least one air controlling agent.
- 25 17. A mortar, concrete, cement or cementitious binder containing the modified acrylic polymer according to anyone of the claims 1 to 9 and 13 in an amount of 0.01 to 10% by weight of the binder, said mortar or concrete having a unit content of binder composed of cement or a mixture of cement and latent hydraulic or inert microscopic powder of 150 to 800 kg/m<sup>3</sup>, preferably of 250 to 650 kg/m<sup>3</sup>.
- 30 18. A mortar, concrete, cement or cementitious binder according to claim 17, wherein the cement is Portland cement, white cement, high alumina cement, and wherein the latent hydraulic or inert microscopic powder is fly ash, slag, silica fume, burnt oil shale, metakaolin or calcium carbonate.
- 35 19. A mortar, concrete, cement or cementitious binder according to claim 17, wherein an admixture of one of claims 14 to 16 is added before, during or after the grinding operation of the cement or the cementitious binder.
20. An aqueous slurry of microscopic powders containing the modified acrylic polymers according to anyone of claims 1 to 9 and 13 in an amount of 0.01 to 10% by weight of the binder.
- 40 21. An aqueous slurry of microscopic powders according to claim 20, **characterized in that** the microscopic powder is calcium carbonate, gypsum or gypsum based solids.



European Patent  
Office

# EUROPEAN SEARCH REPORT

Application Number  
EP 00 10 5911

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
A	WO 93 22358 A (BASF) 11 November 1993 (1993-11-11) * claim 1 *	1	C08F220/04 C04B24/26
A	EP 0 792 850 A (NIPPON SHOKUBAI CO. LTD.) 3 September 1997 (1997-09-03)	/	
A	FR 2 655 875 A (COATEX S.A.) 21 June 1991 (1991-06-21)		
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
			C08F C04B
The present search report has been drawn up for all claims			
Place of search <b>THE HAGUE</b>		Date of completion of the search <b>27 June 2000</b>	Examiner <b>Cauwenberg, C</b>
<p><b>CATEGORY OF CITED DOCUMENTS</b></p> <p>X : particularly relevant if taken alone  Y : particularly relevant if combined with another document of the same category  A : technological background  O : non-written disclosure  P : intermediate document</p> <p>T : theory or principle underlying the invention  E : earlier patent document, but published on, or after the filing date  D : document cited in the application  L : document cited for other reasons  &amp; : member of the same patent family, corresponding document</p>			

**ANNEX TO THE EUROPEAN SEARCH REPORT  
ON EUROPEAN PATENT APPLICATION NO.**

EP 00 10 5911

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report.  
The members are as contained in the European Patent Office EDP file on  
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27-06-2000

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
WO 9322358 A	11-11-1993	DE 4213971 A	04-11-1993
		DE 59303085 D	01-08-1996
		EP 0638098 A	15-02-1995
		ES 2088283 T	01-08-1996
		JP 7505919 T	29-06-1995
EP 792850 A	03-09-1997	CN 1165795 A	26-11-1997
		JP 9286645 A	04-11-1997
		JP 2000063165 A	29-02-2000
		US 5925184 A	20-07-1999
		JP 9286646 A	04-11-1997
FR 2655875 A	21-06-1991	AT 102961 T	15-04-1994
		AU 7067791 A	18-07-1991
		CA 2070636 A	21-06-1991
		DE 69007495 D	21-04-1994
		DE 69007495 T	20-10-1994
		EP 0506808 A	07-10-1992
		FI 922809 A	17-06-1992
		WO 9109067 A	27-06-1991
		JP 5502692 T	13-05-1993
		NO 922398 A	18-06-1992
		ZA 9010185 A	30-10-1991